beyond doubt that the strong bands are associated only with the mixed-valence ions. However, the electronic spectrum of the  $(\mu$ -pyrazine)-cis-(dinitrogen)nonaamminediosmium 5+ ion is dramatically different, approximating a summation of spectral properties of a tetraammine Os(II) coordinated by dinitrogen and pyrazinium ion and a pentaammine Os(III) pyrazinium center.<sup>6,7</sup> Further, the dinitrogen stretching frequency (2075 cm<sup>-1</sup>) is appropriate for Os(II), and a band with comparable extinction occurs at 1600 cm<sup>-1</sup>. This band has been previously associated with a ring mode of the pyrazine, which is normally IR inactive for  $D_{2h}$  symmetry but becomes allowed when the center of symmetry has been lost, as in localized mixed-valence systems.<sup>18</sup> This 1600-cm<sup>-1</sup> peak is quite weak in  $[(NH_3)_4ClOs(pyz)Os(NH_3)_5]^{4+}$  and is

virtually absent in the symmetrical ( $[(NH_3)_5Os]_2pyz)^{5+}$  ion, in harmony with their delocalized description. A detailed spectra analysis is now in progress<sup>19</sup> to further our understanding of the electronic structures in relation to both a molecular orbital description and the effects of spin-orbital coupling.

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## Additions and Corrections

Polymerized Phosphatidylcholine Vesicles. Synthesis and Characterization [J. Am. Chem. Soc. 1982, 104, 791]. STEVEN L. REGEN,\* ALOK SINGH, GÜNTHER OEHME, and MANINDER SINGH.

Page 793: Figure 5 is in error. The turbidity of both polymerized and nonpolymerized vesicles of 5 decreases approximately 30% on going from 0% to 25% ethanol. When 45% ethanol in water is used, nonpolymerized vesicles are completely destroyed (no apparent absorbance at 400 nm) and the polymerized dispersion exhibits a turbidity that is 20% higher than that found in pure water. We presume that this increase in turbidity is due to vesicle aggregation. In addition, phosphorus analysis now reveals a small but detectable extractability (CHCl<sub>3</sub>) of lipid monomer, oligomer, and/or polymer from aqueous polymerized dispersions of 5.

Cyclohexaamylose Complexation with Organic Solvent Molecules [J. Am. Chem. Soc. 1982, 104, 6283–6288]. ROBERT I. GELB, LOWELL M. SCHWARTZ,\* MICHAEL RADEOS, ROBERT B. EDMONDS, and DANIEL A. LAUFER.

In our recent paper<sup>1</sup> we regretfully failed to mention a previous publication by Matsui and Mochida<sup>2</sup> which reports studies of some of the same chemical systems. Both papers determine aqueous complexation constants of cyclohexaamylose with ethanol, 2-propanol, 2-methyl-2-propanol, and cyclohexanol. Although the two studies employed entirely different experimental techniques, the sets of equilibrium constant values are in essential agreement.

Reactive Intermediates. Volume 2 [J. Am. Chem. Soc. 1982, 104, 7394].

First column, second line from the bottom, the sentence starting here should read as follows: An idea of the activity in these various fields is given by the fact that the carbene, free-radical, and silylene chapters are by far the longest, whereas that on arynes is less than a fifth as long as the carbene chapter.

Surface Photochemistry: Deviation of the Course of Reaction in Benzoin Ether Photolysis by Adsorption on Silica Gel [J. Am. Chem. Soc. 1982, 104, 6824–6825]. PAUL DE MAYO,\* AKIRA NAKAMURA, PETER W. K. TSANG, and S. KING WONG.\*

Page 6825, Scheme I: Structure 6 should be

Definitive Evidence for Cycloheptatetraene from Dehydrobromination of Bromocycloheptatrienes [J. Am. Chem. Soc. 1982, 104, 7329–7330]. James W. Harris and W. M. Jones.\*

Page 7329: In footnote 3 the date should be 1970 not 1980.

Iron(III)-Porphyrin  $\pi$ -Cation Radical Complexes. Molecular Structures and Magnetic Properties [*J. Am. Chem. Soc.* 1982, 104, 6793]. Georges Buisson, Alain Deronzier, Emile Duee, Pierre Gans, Jean-Claude Marchon,\* and Jean-Rene Regnard.

Page 6795, second column, the second paragraph should read: Pertinent bond lengths in the coordination unit of 3 are as follows: Fe-N<sub>1</sub>, 2.04 (1); Fe-N<sub>2</sub>, 2.05 (2); Fe-O, 2.13 (1) Å. The equatorial bond length values are consistent with one-electron occupancy of the  $d_{x^2-y^2}$  orbital, and therefore they confirm the high-spin Fe(III) state, ...

Kinetics, Thermodynamics, and Mechanism of the Radical Chain Process for Ligand Substitution of Metal Carbonyls [*J. Am. Chem. Soc.* **1983**, *105*, 61–73]. J. W. HERSHBERGER, R. J. KLINGLER, and J. K. KOCHI.\*

Page 63: Entry 18 of Table II for  $(MeC_5H_4)Mn(CO)_3$  should be corrected to— $E_p^{\text{ox}} = 1.15 \text{ V}$  and  $(E_p^{\text{ox}} + E_p^{\text{red}})/2 = 1.10 \text{ V}$ .

Organic Electronic Spectral Data. Volume XVII [J. Am. Chem. Soc. 1982, 104, 7673].

The third editor of this reviewed book should be H. Feuer.

<sup>(18)</sup> Callahan, R. W.; Keene, F. R.; Meyer, T. J.; Salmon, D. J. J. Am. Chem. Soc. 1977, 99, 1064-1073.

<sup>(19)</sup> Lay, P. A.; Magnuson, R. H.; Taube, H., work in progress.

<sup>(20)</sup> It was necessary to use the Kirsten-Dumas method to obtain satisfactory analytical data for nitrogen. (Lay, P. A.; Sargeson, A. M.; Skelton, B. W.; White, A. H. J. Am. Chem. Soc. 1982, 104, 6161-6164.) All microanalyses were performed by the Stanford University Microanalytical Laboratory.

<sup>(21)</sup> Beattie, J. K.; Hush, N. S.; Taylor, P. R. Inorg. Chem. 1976, 15, 992-993.

<sup>(1)</sup> Gelb, R. I.; Schwartz, L. M.; Radeos, M.; Edmonds, R. B.; Laufer, D. A. J. Am. Chem. Soc. 1982, 104, 6283-6288.

<sup>(2)</sup> Matsui, Y.; Mochida, K. Bull. Chem. Soc. Jpn. 1979, 52, 2808-2814.